

# Worksheet Assigning Oxidation Numbers

Name: \_\_\_\_\_

## Worksheet – ASSIGNING OXIDATION NUMBERS

### Oxidation Number Rules:

1. The oxidation number of any pure element is 0.
2. The oxidation number of a monatomic ion equals that charge on the ion.
3. The more electronegative element in a binary compound is assigned the number equal to the charge it would have if it were an ion.
4. The oxidation number of fluorine in a compound is always -1.
5. Oxygen has an oxidation number of -2 unless it is combined with F, in which it is +1 or +2, or it is in peroxide (such as  $\text{H}_2\text{O}_2$  or  $\text{Na}_2\text{O}_2$ ), in which it is -1.
6. Hydrogen is +1, unless combined with a metal, and then it is -1.
7. In compounds, Group 1 is +1, Group 2 is +2, and Aluminum is +3.
8. The sum of the oxidation numbers of all atoms in a neutral compound is 0.
9. The sum of the oxidation numbers in a polyatomic ion equals the charge of the ion.

**Part A:** In the following questions, give the oxidation number of the indicated atoms/ion:

- |   |  |
|---|--|
| 1. N in $\text{N}_2\text{O}_3$ _____        | 16. C in $\text{CH}_4$ _____               |
| 2. S in $\text{H}_2\text{SO}_4$ _____       | 17. Mn in $\text{MnO}_2$ _____             |
| 3. C _____                                  | 18. S in $\text{SO}_3^{2-}$ _____          |
| 4. C in CO _____                            | 19. $\text{Mg}^{2+}$ _____                 |
| 5. Na in NaCl _____                         | 20. $\text{Cl}^-$ _____                    |
| 6. H in $\text{H}_2\text{O}$ _____          | 21. $\text{O}_2$ _____                     |
| 7. Ba in $\text{BaCl}_2$ _____              | 22. $\text{P}_4$ _____                     |
| 8. N in $\text{NO}_2^-$ _____               | 23. Na in $\text{Na}_2\text{S}$ _____      |
| 9. S in $\text{Al}_2\text{S}_3$ _____       | 24. S in $\text{H}_2\text{S}$ _____        |
| 10. S in $\text{HSO}_4^-$ _____             | 25. $\text{Ca}^{2+}$ _____                 |
| 11. Cl in $\text{Fe}(\text{ClO}_2)_3$ _____ | 26. C in $\text{CN}^-$ _____               |
| 12. Fe in $\text{Fe}(\text{ClO}_2)_3$ _____ | 27. H in $\text{OH}^-$ _____               |
| 13. N in $\text{NO}_3^-$ _____              | 28. Mn in $\text{KMnO}_4$ _____            |
| 14. $\text{Cu}^{2+}$ _____                  | 29. I in $\text{Mg}(\text{IO}_3)_2$ _____  |
| 15. $\text{Zn}^{2+}$ _____                  | 30. C in $\text{C}_2\text{O}_4^{2-}$ _____ |

Worksheet assigning oxidation numbers is an essential tool in the study of chemistry, particularly in understanding redox (reduction-oxidation) reactions. Oxidation numbers, also known as oxidation states, provide a way to keep track of electrons in chemical reactions, helping chemists identify which species are oxidized and which are reduced. This article outlines the principles of assigning oxidation numbers, practical applications, and provides a comprehensive worksheet to practice these concepts.

## Understanding Oxidation Numbers

Oxidation numbers are theoretical constructs that help describe the degree of oxidation of an atom in a molecule. They are assigned based on a set of rules that take into account the electronegativity of elements and their bonding characteristics.

# What is an Oxidation Number?

An oxidation number can be defined as:

- A hypothetical charge that an atom would have if all bonds to atoms of different elements were 100% ionic.
- It helps to identify the electron transfer during chemical reactions.

Oxidation numbers can be positive, negative, or zero, and they are crucial in balancing redox reactions, where the total number of electrons lost equals the total number of electrons gained.

## Rules for Assigning Oxidation Numbers

When assigning oxidation numbers, several rules must be followed. Here is a list of the key rules:

1. The oxidation number of an atom in its elemental form is zero.
  - Example:  $O_2$ ,  $N_2$ , and Fe have oxidation numbers of 0.
2. The oxidation number of a monoatomic ion is equal to its charge.
  - Example:  $Na^+$  has an oxidation number of +1, and  $Cl^-$  has an oxidation number of -1.
3. In compounds, the oxidation number of hydrogen is +1, and for oxygen, it is -2.
  - Exception: In peroxides, oxygen has an oxidation number of -1.
4. The sum of oxidation numbers in a neutral compound is zero.
  - Example: In  $H_2O$ , the oxidation number of H is +1 (2 hydrogens) and O is -2, which sums to 0.
5. The sum of oxidation numbers in a polyatomic ion is equal to the charge of the ion.
  - Example: In  $SO_4^{2-}$ , the sum of oxidation numbers is -2.
6. Group 1 elements (alkali metals) always have an oxidation number of +1, and Group 2 elements (alkaline earth metals) always have an oxidation number of +2.
7. Fluorine always has an oxidation number of -1 in compounds.
8. Other halogens (Cl, Br, I) typically have oxidation numbers of -1, unless combined with oxygen or other halogens.

## Practical Applications of Oxidation Numbers

Understanding oxidation numbers is critical in several areas of chemistry, including:

- Balancing redox reactions
- Analyzing reaction mechanisms
- Understanding biological processes such as respiration and photosynthesis
- Industrial processes like metallurgy and electrochemistry

# Balancing Redox Reactions

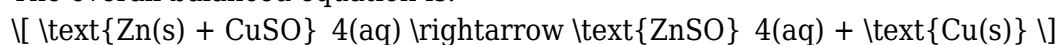
When balancing redox reactions, oxidation numbers help identify which atoms undergo oxidation and which undergo reduction. The following steps outline the typical approach:

1. Assign oxidation numbers to all atoms in the reaction.
2. Identify the species that are oxidized and reduced.
3. Balance the changes in oxidation numbers by adjusting coefficients.
4. Ensure mass and charge balance in the final equation.

For example, in the reaction of zinc with copper sulfate:

- $\text{Zn(s)} \rightarrow \text{Zn}^{2+}(\text{aq}) + 2\text{e}^-$  (oxidation, Zn goes from 0 to +2)
- $\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu(s)}$  (reduction, Cu goes from +2 to 0)

The overall balanced equation is:



## Oxidation Numbers in Biological Processes

In biochemistry, oxidation numbers play a vital role in understanding metabolic pathways. For instance, during cellular respiration, glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ ) is oxidized to carbon dioxide ( $\text{CO}_2$ ), and oxygen ( $\text{O}_2$ ) is reduced to water ( $\text{H}_2\text{O}$ ).

1. Oxidation of Glucose:
  - The oxidation number of carbon in glucose decreases as it is converted to  $\text{CO}_2$ .
2. Reduction of Oxygen:
  - The oxidation number of oxygen decreases as it is converted to water.

These changes are crucial for energy transfer in cells, highlighting the importance of oxidation states in biochemistry.

## Industrial Applications

In industry, oxidation numbers are significant in processes such as:

- Electrolysis: Understanding oxidation states helps in the extraction of metals from ores.
- Corrosion: The oxidation of metals plays a critical role in corrosion processes, which can be mitigated through protective coatings.
- Synthesis of Chemicals: Knowledge of oxidation states aids in the synthesis of various compounds, ensuring desired products are formed.

## Worksheet for Assigning Oxidation Numbers

To reinforce the understanding of assigning oxidation numbers, a worksheet can be an effective tool.

Below is a sample worksheet that can be used for practice.

### Worksheet: Assigning Oxidation Numbers

1. Assign oxidation numbers to the following compounds:

- a.  $\text{H}_2\text{O}$
- b.  $\text{NaCl}$
- c.  $\text{Ca}(\text{OH})_2$
- d.  $\text{NH}_3$
- e.  $\text{SO}_2$
- f.  $\text{H}_2\text{SO}_4$
- g.  $\text{CH}_4$
- h.  $\text{KMnO}_4$

2. Identify the oxidation and reduction states in the following reactions:

- a.  $2\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Fe} + 3\text{CO}_2$
- b.  $2\text{AgNO}_3 + \text{Cu} \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$
- c.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

3. For each of the following reactions, assign oxidation numbers and identify the oxidizing and reducing agents:

- a.  $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
- b.  $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
- c.  $\text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$

4. Fill in the blank with the correct oxidation number:

- a. In  $\text{H}_2\text{O}$ , H has an oxidation number of \_\_\_\_.
- b. In  $\text{Na}_2\text{SO}_4$ , the oxidation number of S is \_\_\_\_.
- c. In  $\text{H}_2\text{O}_2$ , the oxidation number of O is \_\_\_\_.

5. Create a summary table for the oxidation numbers of common elements:

Element	Oxidation Number(s)
Hydrogen	+1
Oxygen	-2, -1 (in peroxides)
Fluorine	-1
Alkali metals	+1
Alkaline earth metals	+2
Halogens	-1 (unless bonded to O or other halogens)

## Conclusion

In conclusion, the concept of worksheet assigning oxidation numbers serves as a foundational aspect of chemistry that enables a deeper understanding of chemical reactions and their implications. By mastering the rules for assigning oxidation states and applying them to various scenarios, students and professionals alike can enhance their comprehension of both theoretical and practical chemistry. Whether in a classroom setting or a laboratory, the ability to accurately assign oxidation numbers is a critical skill that supports numerous applications across scientific fields.

## Frequently Asked Questions

### What is the purpose of assigning oxidation numbers in a chemical compound?

Assigning oxidation numbers helps to determine the distribution of electrons among atoms in a compound, which is essential for understanding redox reactions and balancing chemical equations.

### How do you assign oxidation numbers to elements in a compound?

To assign oxidation numbers, follow these rules: the oxidation number of an element in its elemental form is 0, for monoatomic ions it equals the charge of the ion, oxygen typically has an oxidation number of -2, and hydrogen usually has +1 when bonded to nonmetals.

### What is the oxidation number of oxygen in peroxides?

In peroxides, the oxidation number of oxygen is -1, unlike the common -2 oxidation state found in most compounds.

### Can oxidation numbers be fractional, and if so, in what cases?

Yes, oxidation numbers can be fractional in cases such as coordination complexes and certain organic compounds where resonance structures are involved.

### What are some common mistakes when assigning oxidation numbers?

Common mistakes include incorrectly assigning oxidation numbers to transition metals, overlooking the special cases for oxygen and hydrogen, and neglecting the overall charge of the compound.

### How do you use oxidation numbers to identify redox reactions?

Redox reactions can be identified by tracking changes in oxidation numbers; if the oxidation number of an element increases, it is oxidized, while a decrease indicates reduction.

### What tools or resources can help in practicing oxidation number assignments?

Worksheets, online quizzes, and interactive chemistry platforms are valuable resources for practicing oxidation number assignments and enhancing understanding through examples and exercises.

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