

# Science 8 Density Calculations Worksheet

## SCIENCE 8 – DENSITY CALCULATIONS WORKSHEET

NAME: Key

- 1) A student measures the mass of an  $8 \text{ cm}^3$  block of brown sugar to be  $12.9 \text{ g}$ . What is the density of the brown sugar?

$$m = 12.9 \text{ g} \quad V = 8 \text{ cm}^3 \quad D = \frac{m}{V} \quad D = \frac{12.9 \text{ g}}{8 \text{ cm}^3} \quad \boxed{D = 1.6125 \text{ g/cm}^3}$$

- 2) A chef fills a  $50 \text{ mL}$  container with  $43.5 \text{ g}$  of cooking oil. What is the density of the oil?

$$V = 50 \text{ mL} \quad m = 43.5 \text{ g} \quad D = \frac{m}{V} \quad D = \frac{43.5 \text{ g}}{50 \text{ mL}} \quad \boxed{D = 0.87 \text{ g/mL}}$$

- 3) Calculate the mass of a liquid with a density of  $2.5 \text{ g/mL}$  and a volume of  $15 \text{ mL}$ .

$$D = 2.5 \text{ g/mL} \quad 2.5 \text{ g/mL} = \frac{m}{15 \text{ mL}} \quad m = (2.5)(15) \quad \boxed{m = 37.5 \text{ g}}$$

- 4) Calculate the volume of a liquid with a density of  $5.45 \text{ g/mL}$  and a mass of  $65 \text{ g}$ .

$$D = 5.45 \text{ g/mL} \quad m = 65 \text{ g} \quad 5.45 = \frac{65}{V} \quad V = \frac{65}{5.45} \quad \boxed{V = 11.93 \text{ mL}}$$

- 5) A machine shop worker records the mass of an aluminum cube as  $176 \text{ g}$ . If one side of the cube measures  $4 \text{ cm}$ , what is the density of the aluminum?

$$m = 176 \text{ g} \quad V = L \times W \times H = 4 \times 4 \times 4 = 64 \text{ cm}^3 \quad D = \frac{176 \text{ g}}{64 \text{ cm}^3} \quad \boxed{D = 2.75 \text{ g/cm}^3}$$

- 6) A teacher performing a demonstration finds that a piece of cork displaces  $23.5 \text{ mL}$  of water. The piece of cork has a mass of  $5.7 \text{ g}$ . What is the density of the cork?

$$m = 5.7 \text{ g} \quad V = 23.5 \text{ mL} \quad D = \frac{5.7}{23.5} \quad \boxed{D = 0.24 \text{ g/mL}}$$

- 7) A carver begins work on the following block of granite that weighs  $2700 \text{ g}$ . What is the density of the granite?



$$m = 2700 \text{ g} \quad V = L \times W \times H = 20 \times 10 \times 5 = 1000 \text{ cm}^3 \quad D = \frac{2700}{1000} \quad \boxed{D = 2.7 \text{ g/cm}^3}$$

- 8) A piece of PVC plumbing pipe displaces  $40 \text{ mL}$  when placed into a container of water. If the pipe has a mass of  $18 \text{ g}$ , what is the density of PVC?

$$V = 40 \text{ mL} \quad m = 18 \text{ g} \quad D = \frac{18}{40} \quad \boxed{D = 0.45 \text{ g/mL}}$$

- 9) A solid magnesium flake has a mass of  $1300 \text{ g}$  and a volume of  $743 \text{ cm}^3$ . What is the density of the magnesium?

$$m = 1300 \text{ g} \quad V = 743 \text{ cm}^3 \quad D = \frac{1300}{743} \quad \boxed{D = 1.75 \text{ g/cm}^3}$$

## Science 8 Density Calculations Worksheet

Density is a fundamental concept in science that relates to how much mass is contained in a given volume. In eighth-grade science, students are introduced to the concept of density, allowing them to explore properties of matter and understand how substances interact. A Science 8 density calculations worksheet is a useful tool for reinforcing these concepts through practice problems, real-world applications, and theoretical questions. This article delves into the essential aspects of density, the importance of density calculations in science, and how to effectively use a worksheet to enhance learning.

# Understanding Density

Density is defined as the mass of an object divided by its volume. The formula for calculating density is:

$$\text{Density (D)} = \frac{\text{Mass (m)}}{\text{Volume (V)}}$$

Where:

- Density (D) is expressed in units such as grams per cubic centimeter (g/cm<sup>3</sup>) or kilograms per cubic meter (kg/m<sup>3</sup>).
- Mass (m) is measured in grams (g) or kilograms (kg).
- Volume (V) can be expressed in milliliters (mL), liters (L), cubic centimeters (cm<sup>3</sup>), or cubic meters (m<sup>3</sup>).

Understanding density is crucial for various scientific fields, including chemistry, physics, and engineering. It helps in identifying substances, predicting behavior in mixtures, and determining buoyancy.

## The Importance of Density in Science

- Identifying Substances:** Density can serve as a unique fingerprint for substances. For example, if the density of an unknown liquid matches that of water (1 g/cm<sup>3</sup>), it may be water or a similar liquid.
- Buoyancy and Floating:** Objects less dense than the fluid they are placed in will float, while denser objects will sink. This principle is crucial in understanding why ships float and how submarines dive.
- Material Properties:** In material science, density helps in determining the strength, conductivity, and thermal properties of materials.
- Chemical Reactions:** Density plays a role in predicting how substances will react, particularly in solutions and mixtures.

## Components of a Density Calculations Worksheet

A well-structured density calculations worksheet will typically include a variety of problems and exercises that challenge students to apply their understanding of density. Here are some common components:

### 1. Definition and Formula Section

This section provides students with the definition of density and the formula used for calculations. It may include examples of different units of measurement and conversions between them.

## 2. Sample Problems

Sample problems allow students to practice calculating density using given mass and volume. For instance:

- Example 1: Calculate the density of an object with a mass of 200 grams and a volume of 50 cm<sup>3</sup>.

$$\begin{aligned} & \left[ \right. \\ D &= \frac{m}{V} = \frac{200 \text{ g}}{50 \text{ cm}^3} = 4 \text{ g/cm}^3 \\ & \left. \right] \end{aligned}$$

- Example 2: An object has a density of 3 g/cm<sup>3</sup> and a volume of 10 cm<sup>3</sup>. What is its mass?

$$\begin{aligned} & \left[ \right. \\ m &= D \times V = 3 \text{ g/cm}^3 \times 10 \text{ cm}^3 = 30 \text{ g} \\ & \left. \right] \end{aligned}$$

## 3. Real-World Applications

This section can include scenarios where students need to apply their knowledge of density to solve real-life problems, such as:

- Determining whether an object will float in water.
- Calculating the density of a mixture of liquids.
- Comparing the densities of different materials to decide which is suitable for a specific application.

## 4. Practice Exercises

A significant portion of the worksheet should be dedicated to practice exercises. These can include:

- Calculating Density: Given mass and volume, students calculate density.
- Finding Mass or Volume: Students may be given density and one other variable (mass or volume) to find the missing component.
- Multi-step Problems: Complex problems that require multiple calculations, such as determining the density of a compound made of two substances.
- True/False Statements: Students assess statements about density, buoyancy, and material properties to check their understanding.

## 5. Reflection Questions

At the end of the worksheet, reflective questions can help students think critically about what they learned. Sample questions include:

- How does temperature affect the density of liquids?

- Why do some objects float while others sink when placed in water?
- Can two substances with different densities mix? Why or why not?

## **Tips for Completing a Density Calculations Worksheet**

To maximize learning outcomes when working on a density calculations worksheet, students can follow these tips:

1. **Read Instructions Carefully:** Before attempting any problems, students should ensure they understand what is being asked.
2. **Use Units Consistently:** Ensure that mass and volume are in compatible units to avoid confusion during calculations.
3. **Show Work:** Writing down the steps taken to solve a problem helps in understanding the process and is crucial for checking work.
4. **Check Answers:** After completing the worksheet, students should review their answers to identify any mistakes and understand where they went wrong.
5. **Ask Questions:** If students encounter difficulties, they should not hesitate to seek help from teachers or classmates.

## **Conclusion**

The Science 8 density calculations worksheet is an essential educational resource that aids students in grasping the concept of density and its applications. By engaging with various problems and exercises, students can solidify their understanding of mass, volume, and the relationship between the two. Through practice, they develop critical thinking skills, improve problem-solving abilities, and learn to apply scientific principles to real-world situations. Understanding density is not just a fundamental skill in science; it is a key to unlocking a deeper knowledge of the physical world. With the right tools, such as a well-designed worksheet, students can make significant strides in their scientific education.

## **Frequently Asked Questions**

### **What is density and how is it calculated?**

Density is defined as mass per unit volume. It is calculated using the formula:  $\text{Density} = \text{Mass} / \text{Volume}$ .

### **What units are commonly used for measuring density?**

Common units for density include grams per cubic centimeter ( $\text{g/cm}^3$ ) for solids, kilograms per cubic meter ( $\text{kg/m}^3$ ) for gases, and grams per liter ( $\text{g/L}$ ) for liquids.

## **How do you determine the density of an irregular object?**

To find the density of an irregular object, measure its mass using a balance and its volume by water displacement in a graduated cylinder, then apply the density formula.

## **What is the significance of density in identifying materials?**

Density is a physical property that helps identify materials because different substances have unique densities, allowing for differentiation based on mass and volume.

## **Can temperature affect the density of a substance?**

Yes, temperature can affect density. Generally, as temperature increases, the volume of a substance expands, which decreases its density.

## **In a density calculation worksheet, what types of problems can students expect?**

Students can expect problems involving direct calculations of density, comparisons of densities of different substances, and applications of density in real-world scenarios.

## **What is a common mistake students make in density calculations?**

A common mistake is incorrectly measuring either mass or volume, leading to inaccurate density calculations; it's crucial to ensure measurements are taken accurately.

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