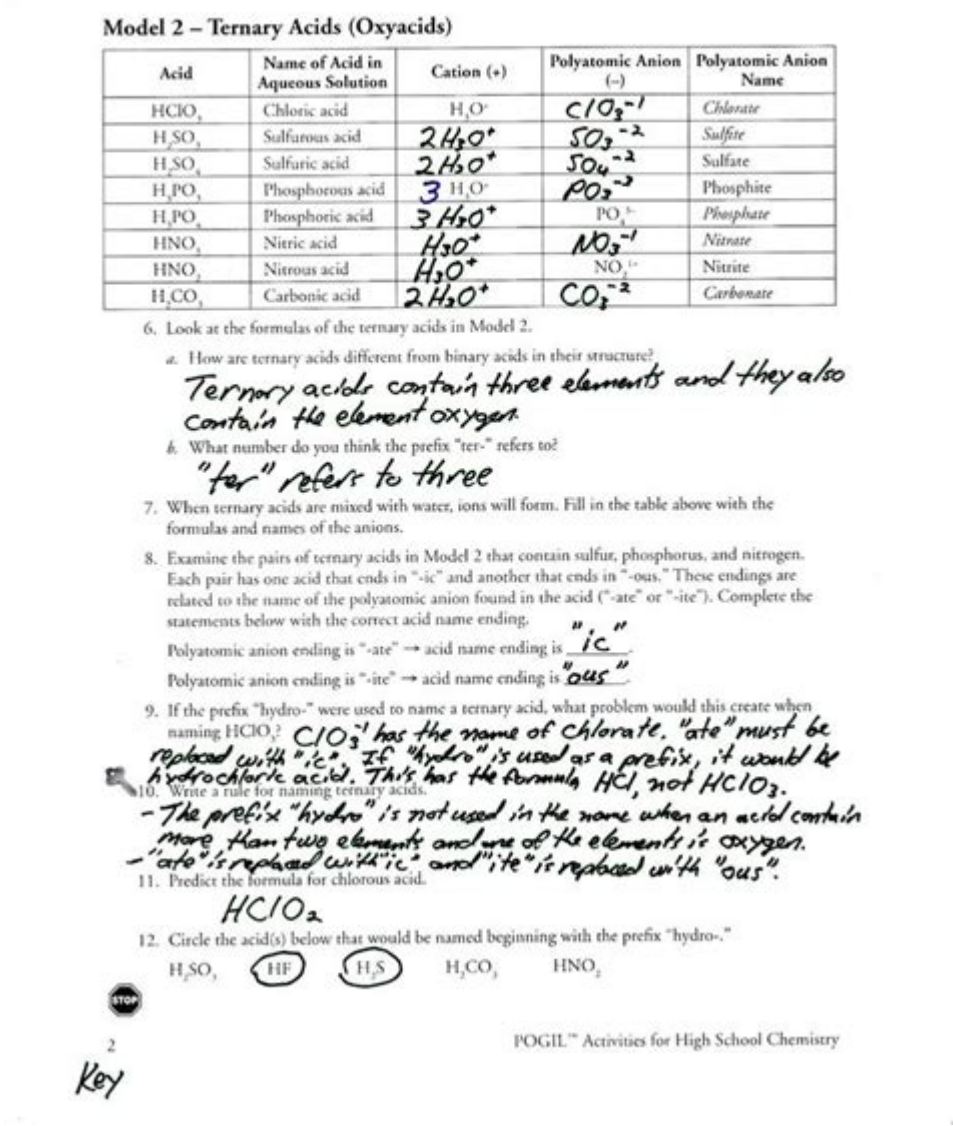
Pogil Naming Acids And Answers



Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (–)	Polyatomic Anion Name
$HClO_2$	Chloric acid	H_3O^+	ClO_3^{-1}	*Chlorate*
H_2SO_3	Sulfurous acid	$2H_3O^+$	SO_3^{-2}	*Sulfite*
H_2SO_4	Sulfuric acid	$2H_3O^+$	SO_4^{-2}	Sulfate
H_3PO_3	Phosphorous acid	$3\ H_3O^+$	PO_3^{-3}	Phosphite
H_3PO_4	Phosphoric acid	$3H_3O^+$	PO_4^{3-}	*Phosphate*
HNO_2	Nitric acid	H_3O^+	NO_3^{-1}	*Nitrate*
HNO_2	Nitrous acid	H_3O^+	NO_2^{1-}	Nitrite
H_2CO_3	Carbonic acid	$2H_3O^+$	CO_3^{-2}	*Carbonate*

6. Look at the formulas of the ternary acids in Model 2.
   a. How are ternary acids different from binary acids in their structure?
   Ternary acids contain three elements and they also contain the element oxygen.
   b. What number do you think the prefix "ter-" refers to?
   "ter" refers to three
7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.
8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.
   Polyatomic anion ending is "-ate" → acid name ending is "ic".
   Polyatomic anion ending is "-ite" → acid name ending is "ous".
9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming $HClO_3$? ClO_3^{-1} has the name of chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl, not $HClO_3$.
10. Write a rule for naming ternary acids.
   - The prefix "hydro" is not used in the name when an acid contain more than two elements and one of the elements is oxygen.
   - "ate" is replaced with "ic" and "ite" is replaced with "ous".
11. Predict the formula for chlorous acid.
   $HClO_2$
12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."
   H_2SO_3 (HF) (H_2S) H_2CO_3 HNO_2

STOP

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Pogil Naming Acids and Answers

Acids are a fundamental category of chemical compounds that play a vital role in various chemical reactions and processes. Understanding how to name acids is essential for students and professionals in the field of chemistry. This article will delve into the Process Oriented Guided Inquiry Learning (POGIL) approach to naming acids, providing a comprehensive overview of the rules, examples, and answers to common questions related to acid nomenclature.

Understanding Acids

Before diving into the nomenclature, it's crucial to understand what acids are. Acids are substances that can donate protons (H^+ ions) in an aqueous solution. They typically have a sour taste and can corrode metals. Acids are classified into two main categories:

1. Binary Acids - Made up of two elements, typically hydrogen and a non-metal.
2. Oxyacids - Contain hydrogen, oxygen, and another element (often a non-metal).

Nomenclature of Binary Acids

Binary acids are named based on the non-metal present in the compound. The naming convention follows these rules:

1. Prefix "Hydro-": Always starts with the prefix "hydro-".
2. Base Name of Non-metal: The base name of the non-metal is used.
3. Suffix "-ic": The name ends with the suffix "-ic".
4. Add "Acid": Finally, the word "acid" is added.

Examples of Binary Acids

- HCl: Hydrochloric acid
- HBr: Hydrobromic acid
- HF: Hydrofluoric acid
- HI: Hydroiodic acid

Nomenclature of Oxyacids

Oxyacids are named differently, depending on the polyatomic ion that forms the acid. The rules for naming oxyacids are as follows:

1. Identify the Polyatomic Ion: Determine the name of the polyatomic ion that contains oxygen.
2. Use the Ion Name: If the polyatomic ion ends in "-ate", the acid name will end in "-ic." If the ion ends in "-ite", the acid name will end in "-ous."
3. Add "Acid": The word "acid" is added at the end of the name.

Examples of Oxyacids

- H_2SO_4: Sulfuric acid (from sulfate)
- H_2SO_3: Sulfurous acid (from sulfite)
- HNO_3: Nitric acid (from nitrate)
- HNO_2: Nitrous acid (from nitrite)
- H_3PO_4: Phosphoric acid (from phosphate)
- H_3PO_3: Phosphorous acid (from phosphite)

Common Questions and Answers about Acid Naming

Understanding the intricacies of acid nomenclature often raises questions. Below are some common inquiries along with their answers.

1. Why do some acids have “-ic” and others have “-ous”?

The distinction arises from the type of polyatomic ion associated with the acid. If an acid is derived from an ion that ends in “-ate”, it takes the “-ic” suffix. Conversely, if it comes from an ion ending in “-ite”, it uses the “-ous” suffix. This system helps in identifying the relationship between the acid and its corresponding polyatomic ion.

2. How do you determine whether to use “hydro-” in binary acid names?

The prefix “hydro-” is only used in binary acids. If the acid consists of only two elements, including hydrogen, you will use the “hydro-” prefix. In contrast, oxyacids do not use this prefix because they contain oxygen as part of their structure.

3. Are there exceptions to these naming rules?

While the rules provided are generally applicable, some acids have historical names that do not follow the systematic nomenclature. For example:

- H_2CO_3: Carbonic acid (derived from carbonate, which follows the rules)
- H_2S: Hydrosulfuric acid (though it can also be referred to as sulfuric acid in certain contexts)

Practice Problems: Naming Acids

To solidify your understanding of naming acids, here are some practice problems. Try to name the following acids before checking the answers.

1. $HClO_4$
2. $H_2C_2O_4$
3. $HBrO_2$
4. H_3AsO_4
5. H_2CO_3

Answers to Practice Problems

1. $HClO_4$: Perchloric acid (from perchlorate)
2. $H_2C_2O_4$: Oxalic acid (historically recognized)
3. $HBrO_2$: Bromous acid (from bromite)
4. H_3AsO_4: Arsenic acid (from arsenate)
5. H_2CO_3: Carbonic acid (from carbonate)

Summary of Key Points

In summary, naming acids is an essential skill in chemistry that requires an understanding of the basic rules for both binary and oxyacids. Here are the key points to remember:

- Binary Acids: Use "hydro-", base name + "ic", and "acid".
- Oxyacids: Name is derived from the polyatomic ion (–ate → –ic, –ite → –ous).
- Compound Identification: Always identify whether the acid is binary or oxyacid before naming.
- Practice: Regular practice with naming acids will enhance your understanding and retention of the rules.

Conclusion

The POGIL approach to learning acid nomenclature emphasizes understanding and applying concepts through guided inquiry and practice. By mastering the naming of acids, students can build a solid foundation for more advanced topics in chemistry. Whether you are a student preparing for exams or a professional looking to refresh your knowledge, this comprehensive guide on naming acids will serve as a valuable resource.

Frequently Asked Questions

What does the acronym POGIL stand for in chemistry education?

POGIL stands for Process Oriented Guided Inquiry Learning.

How are acids named based on the anion present in the compound?

Acids are named based on the anion: if the anion ends in '-ate', the acid name ends in '-ic'; if it ends in '-ite', the acid name ends in '-ous'.

What is the name of HCl when dissolved in water?

HCl dissolved in water is named hydrochloric acid.

How do you name a binary acid?

Binary acids are named by prefixing 'hydro-' to the root of the anion name and adding 'acid', such as H2S being named hydrosulfuric acid.

What is the difference between strong and weak acids in terms of naming?

The naming of acids does not change based on strength; however, strong acids completely dissociate in solution, while weak acids do not.

Give an example of an acid derived from a polyatomic ion and its name.

H2SO4 is derived from the sulfate ion (SO4^2-) and is named sulfuric acid.

What is the significance of the prefix 'per-' in acid naming?

The prefix 'per-' indicates that the acid contains one more oxygen than the '-ate' ion, such as HClO4 being named perchloric acid.

What happens to the naming convention of acids when they contain more than one acidic hydrogen?

Acids with more than one acidic hydrogen are typically named with a prefix indicating the number of hydrogens, such as H3PO4 being named phosphoric acid.

How do you identify the acid formula from its name?

To identify the acid formula from its name, recognize the anion it corresponds to and use the appropriate

prefixes and suffixes to construct the formula.

What is the role of POGIL in understanding acid naming?

POGIL facilitates collaborative learning, allowing students to engage in inquiry-based activities that deepen their understanding of acid naming conventions and their applications.

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