

Lewis Structure Worksheet With Answers

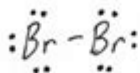
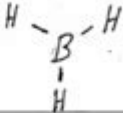
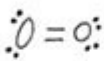
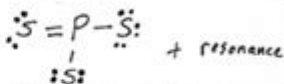
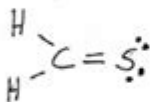
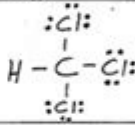
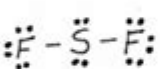
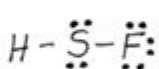
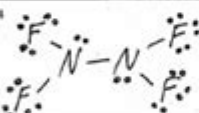
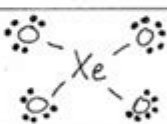
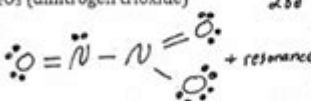
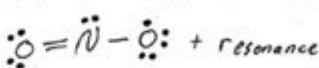
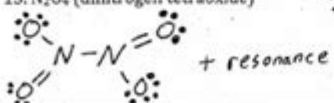
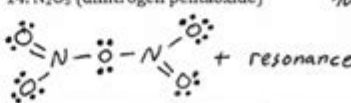
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Lewis Structure Worksheet #2

Key

Draw Lewis Structures for each of the following molecules or ions:

1. Br ₂ 14 	2. BH ₃ 6 
3. O ₂ 12 	4. PS ₃ ⁻ 24 
5. H ₂ CS 12 	6. CHCl ₃ 26 
7. SF ₂ 20 	8. SHF 14 
9. N ₂ F ₄ 38 	10. XeO ₄ 32 
11. N ₂ O ₃ (dinitrogen trioxide)* 28e 	12. NO ₂ (nitrogen dioxide)* 17e 
13. N ₂ O ₄ (dinitrogen tetroxide)* 34 	14. N ₂ O ₅ (dinitrogen pentaoxide)* 40e 

*Problems 11-14 are more challenging and may have more than one correct structure. Please note that if a molecule has an odd number of valence electrons, at least one atom in the molecule will NOT fulfill the octet rule.

Lewis Structure Worksheet #2 - Page 1

Lewis structure worksheet with answers is an essential tool for students and educators alike to understand and visualize the bonding and molecular structure of different chemical compounds. The Lewis structure, also known as the Lewis dot structure, was developed by Gilbert N. Lewis in 1916. It provides a simple way to represent the arrangement of atoms, the distribution of electrons, and the types of bonds between atoms in a molecule. This article will delve into the significance of Lewis structures, how to create them, and provide a worksheet complete with answers for practice.

Understanding Lewis Structures

Lewis structures are graphical representations that show the arrangement of valence electrons in a molecule. They are particularly useful for:

- Predicting molecular geometry
- Understanding chemical reactivity
- Identifying resonance structures
- Determining formal charges

Lewis structures use dots to represent valence electrons and lines to denote covalent bonds between atoms. In the following sections, we will explore how to draw Lewis structures and provide a worksheet with practical examples.

Steps to Draw Lewis Structures

Creating a Lewis structure involves a series of systematic steps:

Step 1: Count Valence Electrons

- Determine the total number of valence electrons from all the atoms in the molecule. This can be done using the periodic table. For example, carbon has four valence electrons, oxygen has six, and hydrogen has one.

Step 2: Identify the Central Atom

- The central atom is usually the least electronegative element (excluding hydrogen) and is typically the atom that can form the most bonds.

Step 3: Connect Atoms with Single Bonds

- Draw single bonds (lines) between the central atom and the surrounding atoms. Each bond represents a pair of shared electrons.

Step 4: Distribute Remaining Electrons

- Place the remaining valence electrons around the outer atoms first to fulfill the octet rule (or duet rule for hydrogen). If there are leftover electrons, place them on the central atom.

Step 5: Form Multiple Bonds if Necessary

- If any atom does not have a complete octet, consider forming double or triple bonds by moving lone pairs from outer atoms to form additional bonds.

Step 6: Check for Formal Charges

- Calculate the formal charge for each atom to ensure stability. A molecule is most stable when the formal charges are minimized and as close to zero as possible.

Worksheet: Lewis Structures Practice

Below is a worksheet with various molecules for which students can practice drawing Lewis structures. After the worksheet, answers will be provided for reference.

Instructions

For each molecule listed below, follow the steps outlined above to draw the Lewis structure.

1. Water (H_2O)
2. Carbon Dioxide (CO_2)
3. Ammonia (NH_3)
4. Methane (CH_4)
5. Sulfur Dioxide (SO_2)
6. Nitrate Ion (NO_3^-)
7. Benzene (C_6H_6)
8. Phosphate Ion (PO_4^{3-})

Answers to the Worksheet

Now, let's go through the answers for the worksheet provided.

1. Water (H_2O)

- Valence Electrons: $\text{O} (6) + 2\text{H} (2) = 8$

- Lewis Structure:

...

H

\

O

/

H

...

- Each hydrogen atom is single-bonded to the oxygen atom.

2. Carbon Dioxide (CO_2)

- Valence Electrons: $\text{C} (4) + 2\text{O} (12) = 16$

- Lewis Structure:

...

$\text{O}=\text{C}=\text{O}$

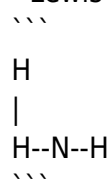
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- The carbon atom forms double bonds with each oxygen atom.

3. Ammonia (NH₃)

- Valence Electrons: N (5) + 3H (3) = 8

- Lewis Structure:

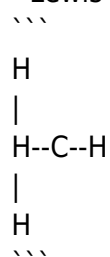


- The nitrogen atom is in the center with three single bonds to hydrogen atoms.

4. Methane (CH₄)

- Valence Electrons: C (4) + 4H (4) = 8

- Lewis Structure:

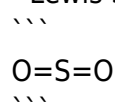


- The carbon atom forms four single bonds with hydrogen atoms.

5. Sulfur Dioxide (SO₂)

- Valence Electrons: S (6) + 2O (12) = 18

- Lewis Structure:

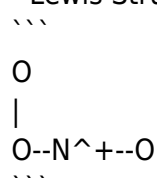


- The sulfur atom forms double bonds with both oxygen atoms.

6. Nitrate Ion (NO₃⁻)

- Valence Electrons: N (5) + 3O (18) + 1 (for the charge) = 24

- Lewis Structure:



- The nitrogen atom has one double bond and one single bond; the overall charge is -1.

7. Benzene (C₆H₆)

- Valence Electrons: 6C (24) + 6H (6) = 30

- Lewis Structure:

```

  \ \
   C
  // \
 C   C
  |   |
  C---C
  \   /

```

- Benzene has alternating double bonds forming a ring structure.

8. Phosphate Ion (PO_4^{3-})

- Valence Electrons: P (5) + 4O (24) + 3 (for the charge) = 32

- Lewis Structure:

```

  O
  ||
O--P--O
  |
  O

```

- The phosphorus atom is at the center with one double bond and three single bonds.

Conclusion

The Lewis structure worksheet with answers serves as a practical exercise for students to master their understanding of molecular geometry and electron distribution in chemical compounds. By following the systematic steps to draw Lewis structures, learners can enhance their comprehension of chemical bonding, which is fundamental in the study of chemistry. Practicing with various molecules, such as water, carbon dioxide, and more complex ions, prepares students not only for academic success but also for future challenges in chemistry and related fields.

Frequently Asked Questions

What is a Lewis structure?

A Lewis structure is a diagram that represents the bonding between atoms of a molecule and the lone pairs of electrons that may exist in the molecule.

Why is it important to use a Lewis structure worksheet?

A Lewis structure worksheet helps students practice drawing Lewis structures, understand molecular geometry, and predict properties of molecules based on their electron configurations.

What information do you typically find on a Lewis structure

worksheet?

A Lewis structure worksheet usually includes spaces to draw structures, questions about the number of valence electrons, and prompts for identifying molecular shapes and polarity.

How do you determine the total number of valence electrons for a molecule?

To determine the total number of valence electrons, add the valence electrons of each atom in the molecule based on their position in the periodic table.

What is a common mistake when drawing Lewis structures?

A common mistake is failing to account for all valence electrons, leading to an incomplete or incorrect representation of the molecule.

How can a Lewis structure worksheet aid in understanding resonance?

A Lewis structure worksheet can include exercises on resonance structures, helping students visualize different ways electrons can be arranged in a molecule without changing the arrangement of atoms.

What is the octet rule, and how does it apply to Lewis structures?

The octet rule states that atoms tend to bond in a way that gives them eight electrons in their valence shell, which is a key principle when drawing Lewis structures.

What are some resources for finding Lewis structure worksheets with answers?

Resources for finding Lewis structure worksheets with answers include educational websites, chemistry textbooks, and online platforms like Khan Academy or Teachers Pay Teachers.

How can you tell if a Lewis structure is correct?

A Lewis structure is correct if it satisfies the octet rule (or duet rule for hydrogen), has the correct number of valence electrons, and accurately represents the molecule's shape and bond types.

What role do formal charges play in Lewis structures?

Formal charges help determine the most stable Lewis structure by indicating the distribution of electrons among atoms, where the ideal structure has the lowest possible formal charges.

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