

Half Life Practice Problems Worksheet

Half-life Practice Worksheet

1. Sodium-24 has a half-life of 15 hours. How much sodium-24 will remain in an 18.0 g sample after 60 hours?
2. After 42 days a 2.0 g sample of phosphorus-32 contains only 0.25 g of the isotope. What is the half-life of phosphorus-32?
3. Polonium-214 has a relatively short half-life of 164 seconds. How many seconds would it take for 8.0 g of this isotope to decay to 0.25 g?
4. How many days does it take for 16 g of palladium-103 to decay to 1.0 g? The half-life of palladium-103 is 17 days.
5. By approximately what factor would the mass of a sample of copper-66 decrease in 51 minutes? The half-life of copper-66 is 5.10 minutes.
6. In 5.49 seconds, 1.20 g of argon-35 decay to leave only 0.15 g. What is the half-life of argon-35?

Half Life Practice Problems Worksheet is an essential tool for students and educators who seek to understand the concept of half-life in nuclear chemistry and radioactive decay. The half-life of a substance is the time required for half of the radioactive atoms in a sample to decay. This concept is pivotal in various fields, including chemistry, physics, and environmental science, as it helps in understanding the behavior of unstable isotopes and their applications. This article will delve into the significance of half-life, provide practice problems to enhance comprehension, and offer detailed solutions to these problems.

Understanding Half-Life

Half-life is a fundamental concept in nuclear chemistry that describes how radioactive

substances decay over time. The half-life of a radioactive element is unique and specific to each isotope. For example, carbon-14 has a half-life of about 5,730 years, while uranium-238 has a half-life of about 4.5 billion years.

The Importance of Half-Life

Understanding half-life is crucial for several reasons:

1. Dating Techniques: Half-life is used in radiometric dating, allowing scientists to determine the age of artifacts, fossils, and rocks.
2. Medical Applications: Radioactive isotopes are used in medical imaging and cancer treatment, making knowledge of half-life essential for determining dosage and safety.
3. Nuclear Waste Management: Knowing the half-lives of radioactive isotopes helps in planning the safe disposal and management of nuclear waste.
4. Environmental Studies: Understanding how long substances remain radioactive aids in assessing environmental contamination and risks.

Half-Life Calculations

Calculating half-life involves understanding exponential decay. The general formula to calculate the remaining quantity of a radioactive substance after a certain time is:

$$N(t) = N_0 \left(\frac{1}{2} \right)^{\frac{t}{t_{1/2}}}$$

Where:

- $N(t)$ = remaining quantity after time t
- N_0 = initial quantity
- t = elapsed time
- $t_{1/2}$ = half-life of the substance

This formula allows us to solve various problems related to half-life.

Practice Problems

Below are several practice problems designed to help students apply their understanding of half-life concepts.

1. Problem 1: A sample of a radioactive isotope has a half-life of 10 years. If you start with 80 grams of the isotope, how much will remain after 30 years?
2. Problem 2: A certain radioactive element has a half-life of 5 days. If you have 200 mg of the substance, how much will remain after 15 days?
3. Problem 3: If a scientist begins with 50 grams of a substance that has a half-life of 4 hours, how much of the substance will remain after 12 hours?

4. Problem 4: A radioactive isotope decays to $\frac{1}{16}$ of its original amount in 24 hours. What is the half-life of this isotope?
5. Problem 5: An isotope has a half-life of 12 years. If the initial amount is 160 grams, how much will remain after 36 years?

Solutions to Practice Problems

Now, let's solve the practice problems step by step.

Solution to Problem 1

Given:

- Initial amount $(N_0 = 80)$ grams
- Half-life $(t_{1/2} = 10)$ years
- Time elapsed $(t = 30)$ years

The number of half-lives that have passed is:

$$\text{Number of half-lives} = \frac{30}{10} = 3$$

Now apply the formula:

$$N(t) = 80 \left(\frac{1}{2} \right)^3 = 80 \left(\frac{1}{8} \right) = 10 \text{ grams}$$

Thus, after 30 years, 10 grams will remain.

Solution to Problem 2

Given:

- Initial amount $(N_0 = 200)$ mg
- Half-life $(t_{1/2} = 5)$ days
- Time elapsed $(t = 15)$ days

Number of half-lives:

$$\text{Number of half-lives} = \frac{15}{5} = 3$$

Now apply the formula:

$$N(t) = 200 \left(\frac{1}{2} \right)^3 = 200 \left(\frac{1}{8} \right) = 25 \text{ mg}$$

Thus, after 15 days, 25 mg will remain.

Solution to Problem 3

Given:

- Initial amount $(N_0 = 50)$ grams
- Half-life $(t_{1/2} = 4)$ hours
- Time elapsed $(t = 12)$ hours

Number of half-lives:

$$\text{Number of half-lives} = \frac{12}{4} = 3$$

Now apply the formula:

$$N(t) = 50 \left(\frac{1}{2} \right)^3 = 50 \left(\frac{1}{8} \right) = 6.25 \text{ grams}$$

Thus, after 12 hours, 6.25 grams will remain.

Solution to Problem 4

Given that the isotope decays to $1/16$ of its original amount in 24 hours, we can set up the equation:

$$\frac{1}{16} = \left(\frac{1}{2} \right)^n$$

Since $\left(\frac{1}{2} \right)^4 = \frac{1}{16}$, we have:

$$n = 4$$

The time for 4 half-lives is 24 hours. Therefore, the half-life $(t_{1/2})$ is:

$$t_{1/2} = \frac{24 \text{ hours}}{4} = 6 \text{ hours}$$

Solution to Problem 5

Given:

- Initial amount $(N_0 = 160)$ grams
- Half-life $(t_{1/2} = 12)$ years
- Time elapsed $(t = 36)$ years

Number of half-lives:

$$\text{Number of half-lives} = \frac{36}{12} = 3$$

Now apply the formula:

$$N(t) = 160 \left(\frac{1}{2} \right)^3 = 160 \left(\frac{1}{8} \right) = 20 \text{ grams}$$

Thus, after 36 years, 20 grams will remain.

Conclusion

The Half Life Practice Problems Worksheet is an invaluable resource for students aiming to grasp the complexities of radioactive decay and half-life calculations. By working through various problems and understanding the underlying concepts, students can develop a robust understanding of how half-life applies across different domains, from scientific research to practical applications in medicine and environmental science. Mastery of these skills not only aids in academic pursuits but also fosters critical thinking and problem-solving abilities essential in scientific disciplines.

Frequently Asked Questions

What is a half-life practice problems worksheet used for?

A half-life practice problems worksheet is used to help students understand and apply the concept of half-life in radioactive decay and other exponential decay processes through practice problems.

How do you calculate the remaining amount of a substance after a certain number of half-lives?

To calculate the remaining amount of a substance after a certain number of half-lives, you can use the formula: $\text{Remaining Amount} = \text{Initial Amount} \times (1/2)^{(\text{number of half-lives})}$.

What types of problems are typically found on a half-life practice problems worksheet?

Typical problems include calculating the amount of a radioactive substance remaining after a specific time period, determining the number of half-lives that have passed, and finding the half-life of a substance given certain data.

Can half-life practice problems be applied outside of chemistry?

Yes, half-life concepts can be applied in various fields such as biology, medicine (e.g., drug elimination), and environmental science (e.g., the decay of pollutants).

What is the importance of mastering half-life calculations?

Mastering half-life calculations is important for understanding decay processes, making predictions in scientific experiments, and applying these concepts in real-world situations such as dating archaeological finds or managing radioactive waste.

Are there any online resources available for half-life practice problems?

Yes, there are many online resources, including educational websites, interactive quizzes, and downloadable worksheets that provide half-life practice problems for students to enhance their learning.

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