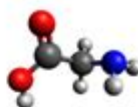


General Chemistry Formula Sheet



Chemistry Formulas



Dilution	$M_1V_1 = M_2V_2$	pK_a & pK_b	$pK_a = -\log[K_a]$
Energy of a photon	$E_{\text{photon}} = hf = hc/\lambda$	pK_a & pK_b	$pK_b = -\log[K_b]$
Pressure	$P = F/A$	pK_a & pK_b	$pK_a + pK_b = 14$
Average Kinetic Energy	$KE_{\text{avg}} = \frac{3}{2}RT$	Neutralization Reaction	$n_A M_A V_A = n_B M_B V_B$
Ideal Gas Law	$PV = nRT$	Buffers	$pH = pK_a + \log[A^-]/[HA]$
Combined Gas Law	$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$	Enthalpy of Formation	$\Delta H_f = \sum n\Delta H_f^{\circ}(\text{product}) - \sum n\Delta H_f^{\circ}(\text{reactants})$
Dalton's Law of Partial Pressures	$P_{\text{total}} = P_A + P_B + \dots$	First Law of Thermodynamics	$\Delta E = q + w$
Dalton's Law of Partial Pressures	$P_A = \chi_A P_{\text{total}}$	Pressure Volume Work	$w = -P\Delta V$
Graham's Law of Effusion	$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$	Calorimetry Thermal Energy (q)	$q = -C_{\text{calorimeter}}\Delta T$
Henry's Law	$P_A = k_H[A]$	Heat Curves & Thermal Energy (q)	$q = mc\Delta T$
Freezing Point Depression	$\Delta T_f = -iK_f m$	Heat Curves & Thermal Energy (q)	$q = m\Delta H_{\text{fusion}}$
Boiling Point Elevation	$\Delta T_b = iK_b m$	Heat Curves & Thermal Energy (q)	$q = m\Delta H_{\text{vaporization}}$
Vapor Pressure Depression (Raoult's Law)	$P_{\text{soln}} = \chi_{\text{solvent}} P_{\text{solvent}}^{\circ}$	Entropy (S)	$\Delta S = \sum nS_{\text{products}} - \sum nS_{\text{reactants}}$
Osmotic Pressure (π)	$\pi = iMRT$	Bond Dissociation Energy	$\Delta H = \sum \Delta H_{\text{reactants}} - \sum \Delta H_{\text{products}}$
Arrhenius Equation	$k = Ae^{-E_a/RT}$	Gibb's Free Energy (ΔG)	$\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$
Equilibrium Constant Expressions	$K_c = \frac{[\text{products}]}{[\text{reactants}]}$	Gibb's Free Energy (ΔG)	$\Delta G = \Delta G^{\circ} + RT\ln Q$
Equilibrium Constant Expressions	$K_{eq} = \frac{k_{\text{forward}}}{k_{\text{reverse}}}$	Gibb's Free Energy (ΔG)	$\Delta G^{\circ} = -RT\ln(K_{eq})$
Equilibrium Constant Expressions	$K_p = \frac{P_{\text{products}}}{P_{\text{reactants}}}$	Standard Cell Potential	$E^{\circ} = E^{\circ}_{\text{reduction}} + E^{\circ}_{\text{oxidation}}$
Reaction Quotient (Q)	$Q = \frac{[\text{products}]}{[\text{reactants}]}$	Standard Cell Potential	$E^{\circ} = E^{\circ}_{\text{cathode}} + E^{\circ}_{\text{anode}}$
Solubility Product Constant (K_{sp})	$K_{sp} = \frac{[\text{products}]}{[\text{reactants}]}$	Nernst Equation	$E_{\text{cell}} = E^{\circ} - \frac{0.0592}{n} \log Q$
pH & pOH	$pH = -\log[H^+]$	Kinetics (always 1st order)	$N = N_0 e^{-kt}$
pH & pOH	$pOH = -\log[OH^-]$	Kinetics (always 1st order)	$\ln N = \ln N_0 - kt$
pH & pOH	$pH + pOH = 14$	Kinetics (always 1st order)	$t_{1/2} = \frac{0.693}{k}$
		Nuclear Binding Energy	$E = \Delta mc^2$

General chemistry formula sheet is an essential tool for students and professionals alike, providing a concise reference for the myriad of equations, constants, and relationships that govern chemical processes. Whether you're preparing for exams, conducting research, or simply looking to refresh your knowledge, a well-organized formula sheet can significantly enhance your understanding and application of chemistry concepts. In this article, we will explore the key components of a general chemistry formula sheet, including fundamental equations, constants, and tips for creating an effective study guide.

Key Components of a General Chemistry Formula Sheet

A comprehensive general chemistry formula sheet typically includes various sections that cover different areas of chemistry. Here are some essential categories to include:

1. Basic Chemical Principles

Understanding the fundamental principles of chemistry is crucial for anyone studying the subject. The following concepts should be included:

- Atomic Structure:
 - Atomic number (Z)
 - Mass number (A)
 - Isotopes
- Periodic Table Trends:
 - Atomic radius
 - Ionization energy
 - Electronegativity
- Mole Concepts:
 - Avogadro's number (6.022×10^{23})
 - Molar mass calculations

2. Chemical Reactions

Chemical reactions are at the heart of chemistry. Including the following equations and types of reactions is important:

- Balancing Chemical Equations:
- General steps for balancing equations
- Example:



- Types of Reactions:
- Synthesis
- Decomposition
- Single replacement
- Double replacement
- Combustion

3. Thermodynamics

Thermodynamics plays a vital role in understanding energy changes during chemical reactions. Key formulas include:

- First Law of Thermodynamics:

$$\Delta U = Q - W$$

where:

- ΔU = change in internal energy
- Q = heat added to the system
- W = work done by the system

- Enthalpy Change:

$$\Delta H = \Delta U + P\Delta V$$

- Gibbs Free Energy:

$$\Delta G = \Delta H - T\Delta S$$

where:

- (T) = temperature in Kelvin
- (ΔS) = change in entropy

4. Stoichiometry

Stoichiometry is essential for understanding quantitative relationships in chemical reactions. Important concepts include:

- Mole Ratios: Derived from balanced equations
- Conversions:
- Moles to grams:

$$\text{grams} = \text{moles} \times \text{molar mass}$$

- Grams to moles:

$$\text{moles} = \frac{\text{grams}}{\text{molar mass}}$$

5. Solutions and Concentrations

In analytical chemistry, understanding solutions and their concentrations is crucial. Include the following formulas:

- Molarity (M):

$$M = \frac{\text{moles of solute}}{\text{liters of solution}}$$

- Dilution Equation:

$$C_1V_1 = C_2V_2$$

where:

- (C_1) and (V_1) = initial concentration and volume
- (C_2) and (V_2) = final concentration and volume

6. Kinetics and Equilibrium

Understanding reaction rates and chemical equilibrium is fundamental in chemistry. Key topics include:

- Rate of Reaction:

$$\text{Rate} = k[A]^m[B]^n$$

where:

- k = rate constant
- $[A]$ and $[B]$ = concentrations of reactants
- m and n = orders of reaction

- Equilibrium Constant (K):

$$K = \frac{\text{products}}{\text{reactants}}$$

Creating Your Own General Chemistry Formula Sheet

A personalized formula sheet can be incredibly beneficial for your studies. Here are some tips on how to create an effective one:

1. Organize by Topic

Organize your formula sheet into clear sections that correspond to different areas of chemistry. This will make it easier to quickly find the information you need when studying or during exams.

2. Use Mnemonics

Incorporating mnemonics can help you remember complex concepts and formulas. For instance, “LEO

says GER” can help you remember that Loss of Electrons is Oxidation, while Gain of Electrons is Reduction.

3. Include Examples

Where applicable, include examples to illustrate the use of formulas. This can help reinforce your understanding of how to apply them in different contexts.

4. Keep it Concise

While it’s important to be comprehensive, try to keep your formula sheet concise. Use bullet points and abbreviations to save space while still conveying the necessary information.

5. Regularly Update It

As you progress in your studies, be sure to update your formula sheet. Add new formulas, concepts, or clarifications as you learn them, ensuring that it remains a valuable resource throughout your chemistry education.

Conclusion

A **general chemistry formula sheet** is an invaluable resource for mastering the various principles, equations, and relationships in chemistry. By organizing your formula sheet into sections covering essential topics like basic chemical principles, reactions, thermodynamics, stoichiometry, solutions, and kinetics, you will create a powerful study aid. Remember to personalize your sheet with examples, mnemonics, and concise notes to enhance your learning experience. With a well-structured formula

sheet, you can approach your chemistry studies with confidence and clarity, paving the way for success in your academic and professional endeavors.

Frequently Asked Questions

What is a general chemistry formula sheet?

A general chemistry formula sheet is a concise reference document that includes essential chemical formulas, constants, and theoretical concepts used in general chemistry.

What key formulas are typically included in a general chemistry formula sheet?

Key formulas often included are the ideal gas law ($PV=nRT$), molarity ($M = \text{moles of solute}/\text{volume of solution}$), and the equations for stoichiometry and thermodynamics.

How can a general chemistry formula sheet help students in exams?

It helps students by providing quick access to important equations and concepts, allowing them to save time during exams and focus on problem-solving.

Are there any specific constants that should be memorized for a general chemistry formula sheet?

Yes, important constants include Avogadro's number (6.022×10^{23}), the ideal gas constant ($0.0821 \text{ L}\cdot\text{atm}/(\text{K}\cdot\text{mol})$), and the molar mass of common elements.

Can a general chemistry formula sheet be customized?

Absolutely! Students can customize their formula sheets based on their course requirements and personal study preferences, including additional formulas or notes.

What is the importance of understanding the concepts behind the formulas on a general chemistry formula sheet?

Understanding the concepts is crucial because it enables students to apply the formulas correctly in different contexts and enhances their problem-solving skills.

Are there online resources available for general chemistry formula sheets?

Yes, many educational websites and platforms provide free downloadable general chemistry formula sheets, along with tutorials and practice problems.

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Unlock the essential concepts of chemistry with our comprehensive general chemistry formula sheet. Perfect for students and enthusiasts! Learn more now!

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