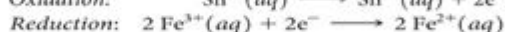
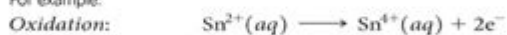


Chapter 20 Review Electrochemistry Section 1

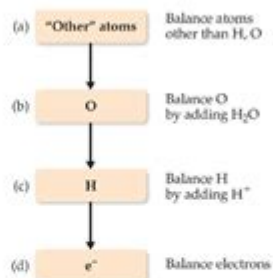
Chapter 20 Electrochemistry

- **Electrochemistry:** study of relationship between electricity and chemical reaction.
- **Oxidation Number:** keep track of which to determine whether oxidation-reduction reaction occurs.
 1. The oxidation number increases for the elements that lose electrons (The element is oxidized).
 2. The oxidation number decreases for the elements that gain electrons (The element is reduced).
- **Half-reaction:** used to balance redox equation

For example:



- **The Half-Reaction Method:**
 1. Divide the equation into one oxidation half-reaction and one reduction half-reaction.
 2. Balance atoms other than O and H. Then, balance O and H using $\text{H}_2\text{O}/\text{H}^{+}$.
 3. Add electrons to balance charges.
 4. Multiply by common factor to make electrons in half-reactions equal.
 5. Add the half-reactions.
 6. Simplify by dividing by common factor or converting H^{+} to OH^{-} if basic.
 7. Double-check atoms and charges balance.
 8. In basic solution, add OH^{-} to each side to "neutralize" the H^{+} in the equation if exists.



- **Voltaic Cell:** a device in which the transfer of electrons takes place through an external pathway. Electrons leave the anode and flow through the wire to the cathode.
 1. The oxidation occurs at the **anode**.
 2. The reduction occurs at the **cathode**.

Chapter 20 Review: Electrochemistry Section 1

Electrochemistry is a fascinating branch of chemistry that deals with the relationship between electricity and chemical reactions. In Chapter 20, Section 1, we delve into the fundamental concepts of electrochemistry, exploring key principles, definitions, and applications. This section sets the stage for understanding more complex topics in electrochemistry, including redox reactions, electrochemical cells, and the practical applications of these principles in real-world scenarios. This comprehensive

review will cover the essential aspects of electrochemistry introduced in this chapter.

Understanding Electrochemistry

Electrochemistry combines chemistry and electrical energy, exploring how chemical reactions can produce electrical energy and vice versa. This section introduces several key concepts that are foundational to the study of electrochemical systems.

Key Definitions

Before delving into the core concepts, it is essential to understand some critical terms used in electrochemistry:

- Oxidation: The loss of electrons from a substance during a chemical reaction.
- Reduction: The gain of electrons by a substance in a chemical reaction.
- Redox Reaction: A chemical reaction that involves the transfer of electrons between two species, encompassing both oxidation and reduction processes.
- Electrochemical Cell: A device that converts chemical energy into electrical energy or vice versa.
- Anode: The electrode where oxidation occurs in an electrochemical cell.
- Cathode: The electrode where reduction occurs in an electrochemical cell.

Redox Reactions

Redox reactions are central to the study of electrochemistry. Understanding how electrons are transferred between reactants is crucial for analyzing electrochemical processes.

The Oxidation States

Oxidation states help track the movement of electrons in a reaction. Each element in a compound has an assigned oxidation state, which reflects its ability to lose or gain electrons. Here are some rules for determining oxidation states:

1. The oxidation state of an element in its standard state is zero.
2. For monoatomic ions, the oxidation state is equal to the charge of the ion.
3. In compounds, hydrogen usually has an oxidation state of +1, while oxygen typically has an oxidation state of -2.

4. The sum of oxidation states in a neutral compound must equal zero, while in a polyatomic ion, it must equal the ion's charge.

Identifying Oxidation and Reduction

To identify which species is oxidized and which is reduced, follow these steps:

1. Assign oxidation states to all elements in the reaction.
2. Determine which elements' oxidation states increase (oxidation) and which decrease (reduction).
3. Balance the overall reaction to ensure that the number of electrons lost equals the number of electrons gained.

Electrochemical Cells

Electrochemical cells are devices that harness the energy from redox reactions to generate electricity. They can be classified into two main types: galvanic (or voltaic) cells and electrolytic cells.

Galvanic Cells

Galvanic cells convert chemical energy into electrical energy through spontaneous redox reactions. Key components of a galvanic cell include:

- **Electrodes:** Two electrodes are immersed in an electrolyte solution, with the anode and cathode facilitating oxidation and reduction, respectively.
- **Salt Bridge:** A salt bridge maintains electrical neutrality in the half-cells by allowing the movement of ions.
- **External Circuit:** This allows electrons to flow from the anode to the cathode, producing an electric current.

Electrolytic Cells

Electrolytic cells, in contrast, use electrical energy to drive non-spontaneous chemical reactions. They are commonly used for processes such as electrolysis, where electrical energy is used to decompose compounds.

Key characteristics of electrolytic cells include:

- **Power Source:** An external power source provides the necessary energy to drive the reaction.
- **Electrode Roles:** The anode is positive, and the cathode is negative,

opposite to galvanic cells.

Standard Electrode Potentials

Standard electrode potentials (E°) are critical for predicting the direction of redox reactions in electrochemical cells. They provide a measure of the tendency of a species to be reduced.

Reference Electrode

The standard hydrogen electrode (SHE) is the reference electrode against which all other electrode potentials are measured. Its potential is defined as 0.00 volts. The potential of other electrodes can be determined relative to the SHE.

Calculating Cell Potential

To calculate the cell potential (E°_{cell}) for a galvanic cell:

1. Identify half-reactions: Write the oxidation and reduction half-reactions.
2. Find standard potentials: Look up the standard reduction potentials for each half-reaction.
3. Calculate E°_{cell} : Use the formula:
$$E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}}$$
4. A positive E°_{cell} indicates a spontaneous reaction, while a negative value suggests a non-spontaneous reaction.

Applications of Electrochemistry

Electrochemistry has numerous practical applications in various fields, from energy production to industrial processes. Here are some notable examples:

Battery Technology

Batteries are electrochemical cells that store and release electrical energy through redox reactions. Common types of batteries include:

- Lead-Acid Batteries: Used in vehicles and for backup power.
- Lithium-Ion Batteries: Widely used in portable electronics and electric

vehicles due to their high energy density and rechargeability.

Corrosion Prevention

Understanding electrochemical processes is crucial in preventing corrosion, which can lead to significant material degradation. Techniques such as cathodic protection and galvanization are employed to protect metals from corrosion.

Electroplating

Electroplating involves depositing a layer of metal onto a surface to enhance appearance, prevent corrosion, or improve conductivity. This process is widely used in the jewelry, automotive, and electronics industries.

Water Treatment

Electrochemical methods are employed in water treatment processes, such as electrolysis, to purify water and remove contaminants effectively.

Conclusion

Chapter 20, Section 1, provides a solid foundation for understanding the principles of electrochemistry. By grasping the concepts of redox reactions, electrochemical cells, and standard electrode potentials, students can appreciate the significance of electrochemistry in both theoretical and practical contexts. The applications of electrochemistry, from battery technology to corrosion prevention, underscore its importance in modern society. As we continue to explore more advanced topics in electrochemistry, the knowledge gained from this section will serve as a critical stepping stone.

Frequently Asked Questions

What is the primary focus of Chapter 20 in electrochemistry?

Chapter 20 primarily focuses on the principles of electrochemical cells, including how they operate, the types of reactions involved, and practical applications.

What is the difference between galvanic cells and electrolytic cells?

Galvanic cells convert chemical energy into electrical energy through spontaneous reactions, while electrolytic cells use electrical energy to drive non-spontaneous chemical reactions.

How does the Nernst equation relate to electrochemistry?

The Nernst equation calculates the cell potential at non-standard conditions, allowing us to understand how concentration changes affect the voltage of electrochemical cells.

What role do electrodes play in electrochemical cells?

Electrodes serve as the sites where oxidation and reduction reactions occur; the anode is where oxidation happens, and the cathode is where reduction takes place.

Why is the concept of standard electrode potential important?

Standard electrode potential provides a reference point for predicting the direction of electrochemical reactions and helps in calculating the overall cell potential.

What are some common applications of electrochemistry in real life?

Common applications include batteries, fuel cells, electroplating, corrosion prevention, and sensors for detecting chemical substances.

What safety precautions should be taken when working with electrochemical cells?

Safety precautions include wearing appropriate personal protective equipment (PPE), ensuring proper ventilation, and handling chemicals according to safety data sheets (SDS).

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